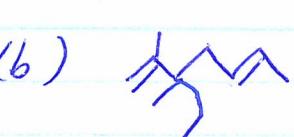
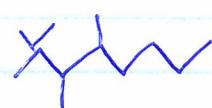


ORGANIC PRACTICE TEST

1. (a)  (b)  (c) 
(d) 

2. both have the formula C_4H_{10} , but they are different shapes

3. non-polar covalent London forces

4. weakest
strongest)
2-methylpropane
butane
2-butyne

- weird shape
- regular shape
- regular shape
but fewer H



(b) slow burning
(c) CO, C

b. uses different boiling points (because each hydrocarbon has different strength of IMFs) to condense each compound at a different point in the column, where it can be removed.

7. A amide
B alkene
C alcohol
D carboxylic acid
E halocarbon.

8. (a) ketone, aldehyde

(b) B, because the =O is more accessible to attract other molecules

9. has a polar end (O-H) and a non-polar end (H-C) So it can dissolve both polar & non-polar compounds